

-6-

In a further aspect of the invention, the concentrates may be used as an infusate in hemodialysis. Consequently, the present invention provides a use of a sterile calcium-free concentrate according to the first embodiment for preparing an infusate for hemofiltration, where said concentrate comprises sodium chloride (NaCl)  $90.72 \pm 9.0$  g/l, magnesium chloride ( $\text{MgCl}_2$ )  ~~$2.05 \pm 0.2$  g/l~~  $0.96 \pm 0.09$  g/l and sodium bicarbonate ( $\text{NaHCO}_3$ )  $28.35 \pm 2.8$  g/l. The present invention also provides a method for hemofiltration comprising administering a sterile dialysis solution comprising Na  $140 \pm 14$  mmol/l, Mg  $0.75 \pm 0.07$  mmol/l, Cl  $116.5 \pm 11$  mmol/l, and  $\text{HCO}_3$   $25.0 \pm 2.5$  mmol/l to a patient in need thereof. The infusate may be prepared by mixing 3000 ml of sterile water to 240 ml of the concentrate.

#### REMARKS

It was discovered during the prosecution in Europe that the original range for magnesium specified as  $2.05 \pm 0.2$  g/l in the disclosure and original claim 1 was an inadvertent error and in fact should clearly be  $0.96 \pm 0.09$  g/l. This correction is requested on the grounds that the inadvertent error would immediately be evident to a person skilled in the art as well as the proper correction required on the following basis.

It would be immediately evident to a skilled person that there must be a direct relationship between the concentration of each component in the dialysis concentrate and the ion components in the diluted dialysis solution.

-7-

In order to achieve a preferred bicarbonate concentration of  $25.0 \pm 2.5$  mmol/L in the dialysis solutions, the dialysis concentrate containing  $28.3 \pm 2.8$  g/l sodium bicarbonate must be diluted by a factor of 13.5. This is entirely consistent with the teaching of the specification at page 13, lines 12-13, where 80ml of bicarbonate concentrate is added to a litre of water to make 1080 ml dialysate, and at page 13, lines 3-6 where one 240 ml unit of concentrate is diluted with 3 litres of water or other suitable diluent.

However, in order for the dialysis solution to have a magnesium concentration of  $0.75 \pm 0.07$  mmol/L, which is expressly stated in the specification at page 2, lines 19-21 and to be variable within only a narrow range, the concentrate must correctly contain magnesium chloride at a concentration of  $0.96 \pm 0.09$  g/l, not  $2.05 \pm 0.2$  g/l.

Thus, not only would this inadvertent error in the calculation of the magnesium chloride concentration in the concentrate have been immediately evident to the skilled person, but also it would have been immediately evident what the correct concentration should be in order to achieve the required magnesium concentration in the diluted dialysis concentrate. Please be advised that this correction was approved in Europe and may be pursued in Canada.

For example to be specific knowing that it is desired to have .75 mmol/l of magnesium in the dialysis solution the following would be the correct simple calculation evident to one skilled in the art to determine the required amount of  $\text{MgCl}_2$  in the dialysis solution. The molecular weight of  $\text{MgCl}_2$  is 95.21 as determined from the attached periodic table from the Merck Index.

-8-

Therefore:  $\frac{.75 \text{ mmol}}{\text{L}} \times 1.08\text{L} = 0.81 \text{ mmol}$

- 0.81 mmol is  $8.1 \times 10^{-4} \text{ mol}$
- but  $\text{MgCl}_2$  has a molecular weight of  $\frac{95.21\text{g}}{\text{mol}}$
- $8.1 \times 10^{-4} \text{ mol} \times 95.21 \text{ g/mol} = 7.71 \times 10^{-2} \text{ g}$
- to convert this value to g/l divide this value by 80ml of concentrate
- $\frac{7.71 \times 10^{-2} \text{ g}}{.08\text{L}} = 0.96375 \text{ g/l}$  rounded to 0.96 plus or minus 10%

Clearly therefore by simple calculations and using the periodic table to determine molecular weights the original value of  $2.05 \pm 10\%$  is clearly incorrect and an inadvertent error.

To verify that this simple calculation was carried out for determination of bicarbonate the same simple calculation is repeated below.

Therefore:  $\frac{25.0 \text{ mmol}}{\text{L}} \times 1.08\text{L} = 27 \text{ mmol}$

27 mmol is  $2.7 \times 10^{-2} \text{ mol}$  but  $\text{NaHCO}_3$  has a molecular weight of 84 g/mol  
 $2.7 \times 10^{-2} \text{ mol} \times 84 \text{ g/mol} = 2.268 \text{ g}$   
 $\frac{2.268 \text{ g}}{.08\text{L}} = 28.35 \text{ g/l} \pm 10\%$  required in concentrate

This value was correctly stated in the disclosure using the same method.

Further to verify that this simple calculation was carried out properly for determination of chloride a similar mass balance calculation was done. These calculations are included in the attached work sheet.

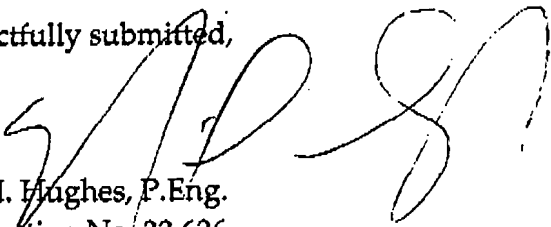
-9-

In determination of chloride in a dialysis solutions one needs to consider the levels of Sodium Chloride and Magnesium Chloride which are both sources of chloride. As determined in the work sheet the numbers specified for NaCl and MgCl<sub>2</sub> in the concentrate and the ions in the dialysis solution are accurate only if the Mg valve is  $.96 \pm 10\%$  and not  $2.05 \pm 10\%$ .

Therefore Applicant respectfully requests entry of this Amendment after Allowance on the record, so that the correction may be included with the printed patent. The Examiner is requested to advise the Applicant of his determination.

Should any questions arise, the Examiner is requested to contact Neil H. Hughes at the office of Applicant's Agents, IVOR M. HUGHES, Barrister & Solicitor, Patent & Trademark Agents at area code (925) 771-6414, at his convenience.

Respectfully submitted,



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NHH/dj

Enclosures:

1. Periodic Table from the Merck Index.
2. Mass Balance Work Sheet.